he va Band of FNO2

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Abstract

are now available for the idea ification of 3NO2, an atnospheric product of CFC and and inproved constants for the ground state. Rotationally resolved infrared line positions consists solely of b type transitions. C.e. 500 infrared and 51 microwave transitions V_4 band associated with the N $_2$ antisynmetric stretching vibration, has been observed in We report the first high resolution infrared spectra of mitryl fluoride $-N D_2$. The intense have been fit to generate rotational constants for the V_{z} - state (V_{0}) ne 750 825 cm egion. The spectrum displays extensive rotational structure and *C oxidation. 793,46684 cm

INTRODUCTION

This paper continues the series of reports from our laboratory on the high resolution spectroscopy of oxyhalide and NO_{λ} compounds (1,2,3) implicated in atmospheric ozone depletion (4). High resolution spectra enable one to determine important structural and bonding properties for these molecules, many of which are unstable. Rotationally resolved line positions can be used to monitor the concentrations of these molecules in atmospheric remote sensing experiments such as ATMOS (5) or in laboratory measurements of reaction kinetics (6).

Nitryl fluoride, FNO₂, has been known as a reactive laboratory compound since the 1930's (7), but has not been the subject of many investigations. It is now known to be one of the principal end products of chlorofluorocarbon oxidation (8). Similar oxidation pathways should exist for the HFC replacement gases, therefore it is important to investigate the spectroscopy and kinetics of FNO₂. We have demonstrated that FNO₂ is the sole product of the F+ NO₂ reaction (9) in contrast to the Cl+ NO₂ reaction where both the ClNO₂ and ClONO structural isomers are formed. Additionally, FNO₂ and its structural isomer FONO are computationally challenging molecules: their bonding, structure and physical properties have been the focus of intense theoretical interest (10).

Microwave spectroscopy on FNO₂ has provided precise ground state rotational constants (11,12,13,14) and a molecular structure (12), but only low resolution infrared spectra have been reported (15,16,17). During work on the $F + NO_2$ reaction system

intense absorption band corresponding to v_e , the NO₂ antisymmetric stretching vibration. investigated under high resolution features due to initiated to obtain high resolution spectra of the 1794 cm⁻¹) bands of ³NO₂. All five regions yielded extensive rotational structure when $(57) \text{ cm}^{-1}$), $v_6 (762 \text{ cm}^{-1})$, $v_2 (82 \text{ cm}^{-1})$ Here we present a full ofational analysis of the DNO isomer we identified infrared V_1 (3.2 cm.) and V_4

EX YERIM ENTA

UV/Visible absorption dedicated OMA (EG&G Princeton A plied Research) to monitor reaction diagnostics via pumping capacity to reduce he esidence time within he .5 meter long cell to signal. This apparatus was designed for the study of transient molecules and has sufficient Ge:Cu detector (Infrared Laboratories) and custom amplifier provided the interferogram optically coupled to a fast-flow/multipass absorption cell (2,3). A liquid heliun cooled nilliseconds while maintaining a total press re of .0 Torr. The system also contains a The spectra were recorded on a Bornem DA8-3 Fourier ansforn spectrometer

Nitry fluoride was not uced by the reaction of Juor ne atoms with N

$$I + NO_2 \xrightarrow{M} FNO_2$$

and helium buffer gas in a highe luorine atoms were produced in a microwave discharge cavity and combined with N nessure pre-reactor region The reaction mixture

flowed from the pre-reactorinto the lower pressure region of the multipass absorption cell for observation. Optimal reaction conditions were determined by monitoring the decay 01 the NO₂ absorpt ion signal in the 300 400" nm region and the concurrent growth of the FNO₂ absorption in the 80 - 240 nm range (18). Under conditions of excess F atoms, any NO₂ introduced into the system was quantitat vely converted into FNO₂. The UV spectra showed no indication of further reaction of the FNO₂ product with F atoms to regenerate NO₂ or to produce F NO. It was hoped that the F + NO₂ system would also produce FONO, by analogy to the reaction of Cl -1 NO₂, but neither the strong 17 15 cm⁻¹ band nor any other absorptions attributable to FONO were observed (9). Survey spectra recorded at i .() cm⁻¹ resolution over the 600 2500 cm⁻¹ range revealed strong absorptions centered at 822, 13.12, and 1794 cm⁻¹ and a sharp, comparatively weak feature near 762 cm⁻¹, in excellent agreement with previous reports of the ENO₂ band positions from low resolution work (16,J7). The reaction conditions could easily be manipulated to saturate the absorptions at 822,1312, and 1 '/(M cm⁻¹. A (Ic.tailed account of the I-4NO₂ reaction may be found in Reference 9.

Spectra of the v_4 band were recorded with a resolution of (),()()5 cm⁻¹ and boxear apodization using a I 600" - 1850 cm⁻¹ band pass filter (OCI J). 128 scans were co-added to obtain the final spectrum. The NO₂ concentration in the reaction mixture was limited so that the strongest features in the spectrum absorbed 1ess than 50%. Infrared line positions were determined using the Kitt Peak peak finder algorithm (19) and calibrated using H₂O transitions (?()) present in the spectrum due to the small amount of residual

water vapor in the optical path. All line positions were assigned an experimental uncertainty of 0.001 cm⁻¹.

SPECTRAL ANALYSIS

FNO₂ is an oblate rotor with C_{2v} symmetry and moderate asymmetry: $\kappa \approx 0.504$ (11,12,13). The presence of two indistinguishable ¹⁶O nuclei, each having zero nuclear spin, requires that the molecular wavefunction follow Bose-Einstein statistics and that the rovibrational component of the wavefunction must have even parity with respect to the molecular a-axis (see Figure 1). The vibrational symmetry of the ground state is a_1 , therefore only even K_a values have non-zero spin-statistical weights. The v_4 =1 state has b_1 vibrational symmetry so that only odd K_a values have non-zero spin-statistical weights in this state. This results in a pure b-type infrared spectrum since all transitions must have $\Delta K_a \approx \pm 1$. The v_4 overview spectrum in Figure 2 exhibits the weak Q branch pattern characteristic of b-type bands.

All spectral calculations were carried out using Pickett's SPET and SPCAT programs (21). Ground state spectroscopic constants were calculated from the 51 microwave transitions reported in References 12, 13, and 14 using Watson's S-reduced Hamiltonian (22)

$$\begin{split} H_{tot} = & \frac{1}{2} (B_{s}^{(S)} + B_{s}^{(S)}) P - [B] - \frac{1}{2} (B_{s}^{(S)} - B_{s}^{(S)}) P_{s}^{2} + \frac{1}{4} (B_{s}^{(S)} - B_{s}^{(S)}) (P_{s}^{2} + P_{s}^{2}) \\ & - D_{J} P^{4} - D_{JK} P^{2} P_{s}^{2} - D_{K} P_{s}^{4} + d_{1} P^{2} (P_{s}^{2} + P_{s}^{2}) + d_{1} (P_{s}^{4} + P_{s}^{4}) \end{split}$$

contained 39 ines with $J \le 15$, $K_e \le 10$ and 10 lines with $-6 \le J \le 29$, ($\le X_a \le 20$. Each in the III representation. The results are presented in Table 1. The microwave data set state constants was sufficiently accurate that it was assumed to predict rotational energy work of Styger et al. (14) were included with 5 kHz uncertainties. This set of ground line was assigned (.05 MHz uncertainty. Two lines from the ultra-high precision FTMW levels for J" = 0-70 with negligible uncertainty

rotational constants for the v_4 : I state were identical to those of the vibrational ground group (23) Initially, 300 transitions between -775 and 1815 cm⁻¹ with $J^{\circ} = 20-40$ were using the interactive Loomis-Wood fitting program LW51 developed by the Giessen there was no central Q-branch. Transition sequences having constant Ka" were identified observed; the P- and R-branches were dominated by clusters of low Ka" transitions and state. This calculation produced a spectrum which was jua itative y similar to the one we assigned and the assignments ver fied using ground state combination differences. Good estimates of C' and (A' +differences A preliminary simulation of the V_4 spectrum was performed assuming the) for the V_4 : 1 state were also determined using combination

Successive fitting iterations involved re-optimizing the spectroscopic constants, assigning additional lines based on the predicted line positions and intensities, and then repeating the process. In this manner, 1539 infrared transitions were incorporated into the final line list. The quantum numbers of the infrared transitions span the range $J^*=0.65$, $K_a^*=0.40$, and $K_c^*=0.65$, thus ensuring an excellent determination of all quartic centrifugal distortion constants. Assigned lines deviating more than 2.5 times the experimental uncertainty from their calculated positions were rejected from the fit. The final constants for the ground state (Table 1) and $V_a=1$ (Table 2) were determined from a simultaneous fit of the combined microwave and infrared data sets. Ground state and $V_a=1$ constants were also determined using Watson's A-reduced Hamiltonian; these values are collected in Table 2.

DISCUSSION

1. The Ground State

There have been several previous investigations of the FNO₂ ground vibrational state using microwave spectroscopy In 1952 Smith and Magnuson (11) identified 8 rotat ional transitions and deduced that the molecule possessed a planar C_{2v} structure. In 1968 Legon and Millen (12) determined a substitution structure for FNO₂ after observing 1.7 rotational transitions in the 17 - 33 GHz region for each Of three isotopomers: F¹⁴N¹⁶O₂, F¹⁵N¹⁶O₂, and I "''N¹³()" The F¹¹N¹⁸O¹⁶O spectrum was particularly interesting since the asymmetric substitution of the oxygen atoms destroyed the C₂

symmetry axis and previously forbidden transitions were observed. This study improved the precision of the A, B, and C rotational constants by three orders of magnitude, but the analysis was constrained to transitions with J < 15 and neglected centrifugal distortion, 1 aterthat year, Mirriet al. ("13) determined the four in-plane centrifugal distortion constants from 33 transitions measured in the 100 - 140 GHz region. [FNO₂ spectroscopy then remained dormant until the recent ultra-high resolution FTMW measurements of Styger et al. (14). which were used to determine the ¹⁴N and ¹⁹F nuclear hyperfine constants.

The rotational constants determined in the present study are collected in Table 1 along with the results of 1 egon and Millen (12) and Mirri et al. (13). The different sets of constants agree very Well. The slightly larger A, B, and C values determined by J egon and Millen probably result from the neglect or centrifugal distortion corrections. The fact that the rotational constants we derived from all 51 of the measured microwave transitions are essentially the same as those derived by M in i et al. supports this collusion.

W C. have converted the centrifugal distortion constants reported by Mirri et al. (13) from the $\tau_{\alpha\beta\beta}$ form into the corresponding S-reduced Hamiltonian constants using Watson's relationships (24,25). The agreement between these values and those derived in the present work is quite good. The IR transitions improve the statistics of the fit without substantially altering the values of the centrifugal distortion constants, even though the IR data set includes transitions up to J'=65, K_a "=40 while the microwave data

quartic centrifugal distortion terms. however, since these constants failed to improve the fit, the lamiltonian was Hamiltonian was also expanded to neluded sextic centrifugal distortion constants; well-deternined and exhibit—arameter/16 error) atios in excess of 800:1. The itting are limited to maxima of J*= 29, K_a*= 20. The final set of ground state constants are all

2. The V₄≈1 State

assignments for each cluster were confir ned using ground state contribution differences. has Ke"=J" and the line strengths dininis) as J"-Ke" increases transitions originate from ground state vibrational levels with the quantum numbers $N_{0,N}$ represents a constant 2"-x value. Beginning with he head of each cluster, the containing several overlapping ransitions. Rotational analysis shows that each cluster of the cluster. The lowest wavenum c feature in each cluster is he most intense, usually cluster consists of a series of lines which converge to a head at the low wavenumber limit spectrum with regular transition clusters occurring a about .42 cm⁻¹ intervals. Each stretching band of 2.N-2, N-22,N 4, N 34,N-6, etc. where N= 2J"-K_c". The strongest line wi hin each cluster The appearance of the v_a band is qualitatively similar to the v_2 NO₂ antisymmetric INO_3 (26). Intense \rightarrow and R-branch transitions dominate ine absolute J

with he transitions $40_{1.40} \leftarrow 4_{-0.4} \ (775.4389 \ cm^{-1}), 39_{1.39} \leftarrow 40_{0.40} \ (1775.8992 \ cm^{-1})$ P-branch clusters for 2J"-K_c" values of 4 40, and 39 beginning

ground and v_4 =1 states ensure that only b-type transitions occur, as mentioned above and $38_{1,38} \leftarrow 39_{0,39}$ (1776,3582 cm⁻¹), respectively. The spin-statistical weights of the resolved. The concentration of transitions near the low wavenumber head of each cluster Since many transitions are "missing" from the spectrum, individual features are easily is clearly evident

derive their intensity from the overlap of two nearly degenerate transitions. lines are due to transitions originating in Ka" J", Ke"~0 levels. Many of these lines and are unrelated to the 2J" K_c"=N clusters. These transitions are characterized by single reduced correlations between all of the V₄: 1 constants. transitions in the fit refined the centrifugal distortion constants D_K/D_{JK} , $\frac{1}{2}$, and d_2 and transitions $19_{17,3} \leftarrow 20_{18,2}$, $19_{17,2} \leftarrow > _{18,3} = 1775.6761 \text{ cm}^{-1}$), and $18_{1,2} \leftarrow$ in the far wings of the spectrum where the 21 - K_c = N clusters are weak (N > 60). These lines having about 0.40 cm⁻¹ spacings. They continue to have appreciable intensity even 19_{18,2} (1776,2942 cm⁻), marked A and — i∋ Figure ≥, are of this type. Including these There are other strong absorptions which run through the P- and R-branch regions

during the early phases of the fitting process. However, the Q-branch line positions and the P- o-R-branch regions. A portion of the Q-branch is shown in Figure 4. The higher intensities were accurately predicted once the rotational constants from fitting the 2- and line density and uncertainty as to absolute J-values prohibited assignment of these lines The Q-branch region between 790 and 1800 cm is much more congested than R-branchregions were available. The Q-branch lines were then incorporated into the fit. This resulted in a further refinements 01 the centrifugal distortion constants and an overall reduction in the 1 σ values for the V4-1 constants.

The fitting residuals are statistically distributed and no detectable perturbations are observed in the spectrum. This is in agreement with our calculations, which predict no other levels within -100" cm 1 of the 1793.466841 cm $^{-1}$ v_{4} origin. The harmonic frequency ω_{4} = 1789 cm $^{-1}$ calculated by Lecatthe CCSD(T)/T Z2P level of theory (10c) compares favor ably with the observed fundamental vibrational frequency. The FNO₂ NO₂ antisymmetric stretching fundamental is significantly higher than the corresponding frequencies observed for HNO₃ (1709.56746 cm $^{-1}$ (26)) and CINO₂ (1683.894022 cm $^{-1}$ (27)), the good agreement between experiment and theory concerning both the trequency and relative intensity of v_{4} suggests that there is nothing unusual about the bonding in FNO₂. The high frequency of this NO_{2} antisymmetric stretching frequency implies that the F atom is withdrawing appreciable electron density from the NO₂ group to form Γ_{5}^{8} NO₂. Recently, Styger et al (14) concluded that in FNO₂ the F atom withdraws 0.36 of an electron from the half-filled NO₂ 6a₁ orbital, thus providing quantitative verification of the electron redistribution.

Given the asymmetry of FNO₂, one would norr nally express the spectroscopic const ants within the framework of Watson's A-reduced Hamiltonian. We have, performed fits using this Hamiltonian and found the quality of the fits comparable to

those obtained using the S-reduced 1 lamiltonian (Table 2). We note, however, that Δ_k is an order of magnitude smaller than expected and that δ_K is an order of magnitude larger than expected. The ratio $\delta_K/\Delta_k = 6.3$ is unphysical and represents optimized fitting parameters rather than reasonable molecular behavior. Therefore, we have included the Δ -reduced constants in Tableonly for the sake of comparison with similar molecules; the S- reduced constants are outpreferred values

CONCLUSIONS

WC. have reported the first rotationally resolved infrared spectrum of FNO2. Over 1500 transitions in the intense v_4 band have been assigned. Well-defined rotational and quartic centrifugal distortion constants have been determined for the ground and v_4 = 1 vibrational states and these constants reproduce the experimental line positions with an rmserror Of 0.00078 cm⁻¹. The high frequency of the NO2 anti-symmetric stretching vibration, $v_0 = 1793.46684$ I cm⁻¹, implies significant redistribution of electron density from the NO2 moiety onto the F atom to form a partially ionic F^8 NO2 structure. However, the small changes in rotational constants upon excitation into the v_4 = 1 state and the ability of high-quality abinitio calculations to reproduce the v_4 vibrational frequency suggest no unusual bonding in FNO2. The v_4 band provides an excellent opportunity for molecule-specific detection of F NO2 since the strongest individual lines coalesce near the lowwavenumber heads of transition clusters in the P-andR-branches to form even more

intense features while still maintaining nearly Doppler-limited linewidths. Analysis of the other vibrational fundamentals is in progress.

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REFERENCES

- M. Birk, R. R. Friedl, E. A. Cohen, H. M. Pickett, and S. P. Sander, J. Chem. Phys. 91, 6588-6597 (1989).
- 2. V. I. Lang, S. P. Sander, and R. R. Friedl, J. Mol. Spectrosc. 132, 89-103 (1988).
- 8 Höjer, R. D. May, and C. E. Miller, J. Quant. Spectrosc. Radiat. Transfer 55, 273-278 (1996).
- Scientific Assessment of Ozone Depletion: 1994, WMO GlobalOzoneResearch and Monitoring Project, Report No. 37, 1994.
- M. R. Gunson, M. M. Abbas, M. C.Abrains, M. Allen, L. R. Brown, T. L. Brown, A.
 Y. Chang, A. Goldman, F. W. Ition, L.L. Lowes, E. Mahieu, G. L. Manney, H. A.

- Michelsen, M. J. Newchurch, C. P. Rinsland, R. J. Salawitch, G. P. Stiller, G. C. Toon, Y. L. Yung, and R. Zander, submitted to *Geophys. Res. Lett.*
- T. S. Dibble, M. M. Maticq, J. J. Szente, and J. S. Francisco, J. Phys. Chem. 99, 17394-17402 (1995).
- R. Schmutzler, Angew. Chem. 80, 466-481 (1968); Angew. Chem. Internat. Edit. 7, 440-455 (1968).
- M. M. Maricq, J. J. Szente, T. S. Dibble, and J. S. Francisco, J. Phys. Chem. 98, 12294-12304 (1994).
- 9. C. E. Miller and S. P. Sander, submitted to J. Phys. Chem.
- 10 a) T. J. Lee and J. E. Rice, J. Chem. Phys. 97, 4223-4232 (1992) and references therein; b) T. S. Dibble and J. S. Francisco, Chem. Phys. Lett. 215, 423-426 (1993); c)
 T. J. Lee, Chem. Phys. Lett. 216, 194-199 (1993).
- 11. D.F. Smith and D. W. Magnuson, Phys. Rev. 87, 226-227 (195?).
- 12. A. C. Legon and D. J. Millen, J. Chem. Soc. A 1968, 1736-1740 (1968).
- 13. A. M. Mirri, G. Cazzoli, and 1. Ferretti, J. Chem. Phys. 49, 2775-2780 (1968).
- 14. C. Styger, B. Gatehouse, N. 1 leineking, W. Jäger, and M. C. L. Gerry, *J.Chem.Soc. FaradayTrans*. 89,1899-19(v (1 993).
- 15. R. E.Dodd, J. A. Rolfe, and L. A. Woodward, 'f rans. Far. Soc. 52, 145-151 (1956).
- 16.11. L.Bernitt, R. 11. Miller, and I.C.Hisatsune, Spectrochim. Acta 23A, ?37-248 (1967).
- 17. K. O. Christie, C. J. Schack, and R.D. Wilson, *Inorg. Chem.* 13, 28 11-2815 (1974).
- 18. A. 1'. Uthman, G. A. Sherman, and (i. A. Takacs, *Spectrochim. Acta* **34A**, **I** 109- **I** 110 (1978).

- 19. 1.. R.Brown, J. S. Margolis, Rt 1. Norton, and B.D. Stedry, Appl. Spectrosc. 37, 287-292 (1983).
- 20 R. A. Toth, J. Opt. Soc. Am. B 8, ? 236-2255 (I 99 I).
- 2 I H. M. Pickett, *J. Mol. Spectrosc.* **148**, 371-377 (199 I).
- 22. J. K. G. Watson, J. Chem. Phys. 48, 451 / 4528 (1968).
- 23. B. 1'. Winnewisser, J. Reinstädtler, K. M. '1'. Yamada, and J. Behrend, J. Mol. Spectrosc. 136, I 2-16 (1989); F. Stroh, Thesis, Justus-J. icbig Universität Giessen (1991).
- 24. W. Gordy and R. L. ('(ink, *Microwa ve Molecular Spectra*, Chapter 8, Wiley Interscience/New York, 1984.
- 25 1.1 ledberg and 1. M. Mills, *J. Mol. Spectrosc.* **160**, 11 '/- 142 (1993).
- 26. A. G. Maki, J. Mol. Spectrosc. 127, 1 04 111 (1988).
- 27 J, Orphal, M. Morillon-Chapey, and G. Guelachvili, 1. Mol. Spectrosc. 165, 315-322 (1994).

LIST 011' TABLES

TABLE 1. FNO₂ Ground Vibrational State Spectroscopic Constants in the S-reduction^a

TABLE 2. Spectroscopic Constants for the 1 NO₂ Ground and v₄: I States^{a,b}

TABLE 1

FNO₂ Ground Vibrational State Spectroscopic Constants in the S-reduction^a

	1.3 7-11	א מד וא	THILL SHIFT	WITHIA CITT
<	(17) / 000 047()	U.44U 348 5 (41)	0.440 241 892 (107)	0.440 341 968 6 (2/4)
; m	0 381 808 7 (27)	0.381 805 8 (27)	0.381 805 757 (76)	0.381 805 842 8 (220)
a C	0.204 109 4 (27)	0.204 107 5 (27)	0.204 107 646 (64)	0.204 107 703 (23)
7.0./.0		5.495 8 ⁵	5.494 23 (454)	5.496 27 (127)
ر ا آر		- x 0 3 x 2 x	(1:8) 60 6868-	-8.98672 (147)
)		5 C 7 C 7 S 8 S 7 S 8 S 7 S 8 S 7 S 8 S 7 S 8 S 7 S 8 S 7 S 8 S 7 S 8 S 8	4.060 38 (604)	4.057.57 (220)
		5.809.35 20.809	-0.800 19 (230)	(0.801, 781, (4.50))
da/10 ⁻³		0.278 423	0.276 39 (163)	0.277 46 (317)
42/10 T	ATA NAME	MA CE	MALE 7 MA 67	49 WW. 2 FTWW. 1511 IR

All values expressed in cm.1. Numbers in parentheses after each constant reflect the 15 error

in units of the last digit.

⁵The quartic centrifugal distortion constants given in Ref. 13. are (in MHz):

$$\tau_{aaa} = -0.039812 \pm 0.000256$$

$$\tau_{bbb} = -0.078638 \pm 0.000087$$

$$\tau_{achb} = 0.012452 \pm 0.000375$$

$$\tau_{abab} = -0.048868 \pm 0.000115$$

These values have been converted into the spectroscopic constants for the S-reduction using the

appropriate relationships (24,25).

TABLE 2 $\label{eq:spectroscopic} \textbf{Spectroscopic Constants for the FNO}_2 \textbf{ Ground and } \nu_4 = 1 States^{a.b}$

S-reduced F	Hamiltonian	A-reduced Hamiltonian			
A"	0.440 347 968 6 (274)	A"	0.440 348 811 2 (327)		
В"	0.381 805 842 8 (220)	В"	0.381 804 456 2 (310)		
C.,	0.204 107 703 (23)	C"	0.204 107 964 (26)		
D ₂ "/10 ⁻⁷	5.496 27 (127)	Δ <u>ι</u> "/10 ⁻⁷	4.928 3 (47)		
D _{!K} "/10 ⁻⁷	-8.986 72 (147)	Δ_{IK} "/10 ⁻⁷	-5.617 6 (237)		
D _K "/10-7	4.057 57 (220)	Δ _K "/"0 ⁻⁷	1.249 (20)		
d-"/10 ⁻⁷	-0.801 781 (450)	8.1710 ⁻⁷	0.808 19 (260)		
$d_2^{-1}/10^{-7}$	0.277 46 (317)	δ _K "/ 'O"	-7.920 8 (494)		
AA	-0.002 843 52 (45)	A`-A``	-0.002 843 35 (51)		
B3	-0.000 186 62 (44)	B'-b	-0.000 186 62 (56)		
CC.,	-0.000 654 576 (153)	C. C.,	-0.000 654 496 (157)		
(D ₁ ''- D ₂ ''')/10 ⁻⁹	-2.371 (129)	(Δ·'-Δ·")/10 ⁻⁹	1.194 (273)		
(D;x'-D;x'')/'0 ⁻⁹	3.956 (237)	$(\Delta_{iK} - \Delta_{iK}) / i 0^{-9}$	-17.15 (102)		
(Dx'-Dx'')/10 ⁻⁹	-0.928 (149)	$(\Delta_{\mathbf{K}}^{\mathbf{Y}} - \Delta_{\mathbf{K}}^{\mathbf{Y}})/1 0^{-9}$	16.67 (81)		
$(d_1 - d_1)^{10^{-9}}$	2.381 (159)	$(\delta_{!} - \delta_{!}) / 10^{-9}$	-2.405 (195)		
$(d_2' - d_2'')/10^{-9}$	-1.712 (75)	(δ _K `-δ _K ``)/10 ⁻⁹	17.28 (217)		
Vç	1793.466 841 (70)	∇c	1793.466 816 (75)		
rms error	0.000 78	rms error	0.000 79		

^aAll values expressed in cm⁻¹. Numbers in parentheses after each constant reflect the locarror in units of the last digit.

⁵All constants determined from the data set consisting of 49 MW, 2 FTMW and 1511IR transitions

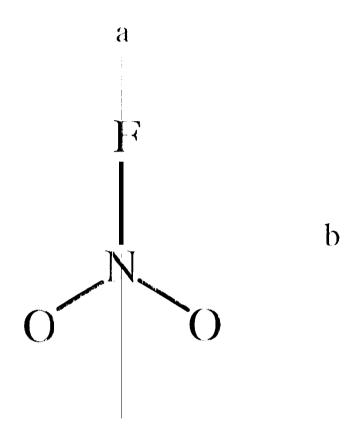
FIGURE CA' TONS

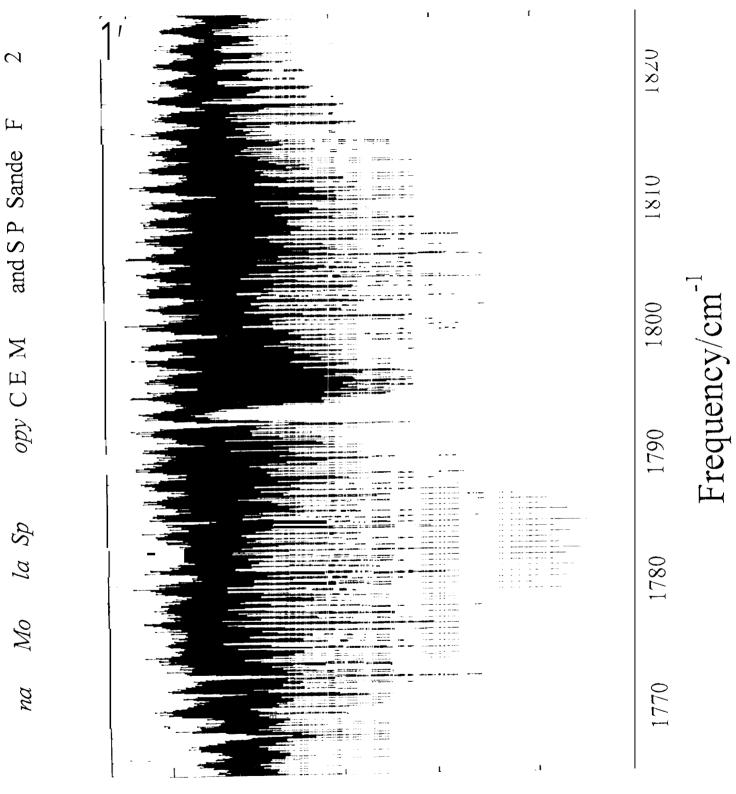
(Rc ' The e-axis is perpendicular to the page. $R_{\rm NJ}=146.7~{\rm pm}, R_{\rm N}$ Figure 1. The molecular structure of FN 2, showr supercaposed on the molecular axes. 17.98 pm, 0_{0NO} =

Figure 2. An overview of the v_4 spectrum recorded at 0.005 cm⁻¹ resolution.

V₄ P-branch. See the text for details. Figure 3. An expanded view of several 21'-Ke": N clusters and two high Ka" lines in the

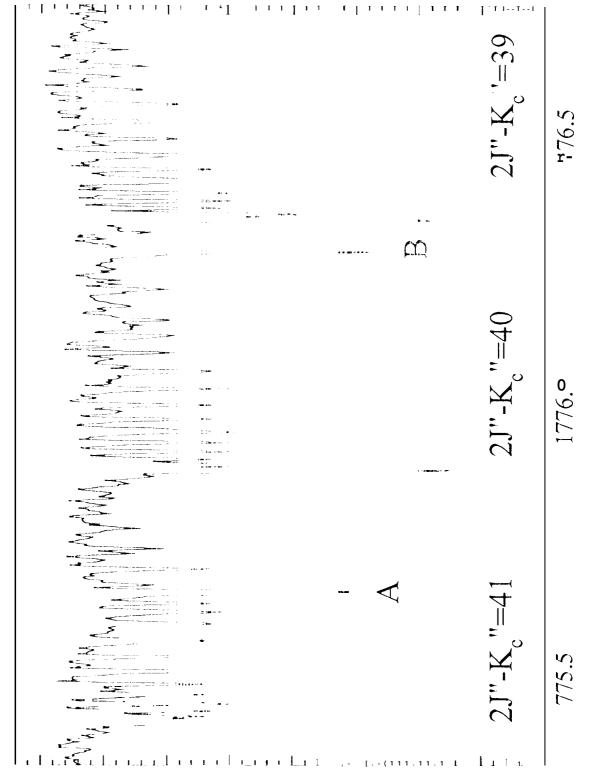
Figure 4. A detailed view of the V₄ spectrum near the vibrational origin.





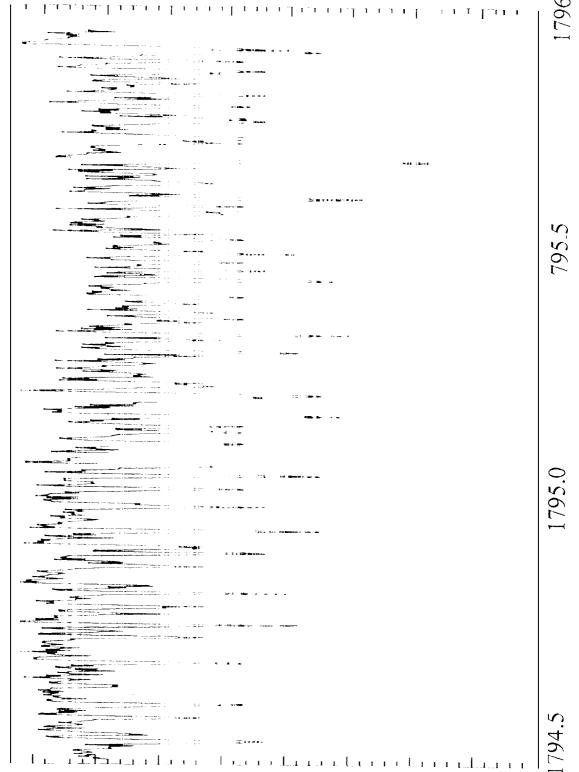
Absorption/arb. units

Absorption/arb. units



Frequency/cm⁻¹

Absorption/arb. units



Frequency/cm⁻¹

L ST OF SYMBOLS

V 0	Vaaaa, Vhbbb, Vaabb, Vabab	$ au_{0\%}$	$\Lambda_{K}, \Lambda_{J}, \Lambda_{JK}, \delta_{J}, \delta_{K}$	J_K J , J_{JK} , d_1 , d_2	^', ' , n.''	A", B", C"	Symbol A, B, C
Origin of a vibrational transition	Axis specific quartic centrifugal distortion constants	Distortion constant for the anular momentum operator $\mathbf{P}_a\mathbf{P}_b\mathbf{P}_s$	Quar ic centrifual distortion constants in the Arceduction	Quartic centrifual distortion constants in the Sreduction	Principal rotational constants for the upper state in a specified transition	Principal rotational constants for the lower state in a specified transition	Definition Principal otational constants